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Answers logical
answer. • The unit of
the known (FSW 3 HP
2) cancels. • The
answer has the correct
unit (W). 3 Evaluate
Does the result make
sense? Chapter 12
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Answer Key Answer:

4.93×10^{-5} L or 49.3

μL In Example 12.2.1

and Example 12.2.2,

the identity of the

limiting reactant has

been apparent:

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[Au(CN)₂]⁻, LaCl₃,
ethanol, and para

-nitrophenol. When the
limiting reactant is

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W/ answers Website
Upload Big Numbers
and Chemistry At the
most fundamental
level, the chemist
needs a unit that
describes a very large
quantity. One of the

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most well-known numbers in the study of chemistry is number of units in a mole. The number of

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This is a problem. But,

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begin reading, it will be
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Grams to Moles or
Moles to Grams:
Answer the following
questions. Be sure to
show all your work and

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your units. (pg.

390-398) 4. The

decomposition of

potassium chlorate,

KClO_3 , is used as a

source of oxygen in the

laboratory.

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relationships between
reactants and products
in a chemical rxn. Mole
Ratio

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- Unit 8- Stoichiometry
VOCABULARY

Assignment:
stoichiometry
percentage yield mole
ratio mass-mass
problem limiting
reagent excess reagent
OBJECTIVES: Be able to
do stoichiometry
problems (mass-mass
problems). Be able to
calculate the limiting
reagent for a given
chemical reaction.

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Chemical Reactions
and Stoichiometry.

Given the equation $3A + B \rightarrow C + D$, you react
1 mole of A with 3

moles of B. True or
false: B; is the limiting
reactant because you
have fewer moles of B
than A. Answer: False.

NaHCO_3 ; is the active
ingredient in baking
soda. How many grams
of oxygen are in 0.35 g

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of. NaHCO₃? a) 0.012

g. b) 0.022 g. c) 0...

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equations 3.

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calculations mass-mass
and volume-volume 4.

Limiting reagent 5.

Percent yield Answer

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