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## **Electron Configuration Practice Answer Key**

Electron Configuration Practice Chemistry Name : Due Date: B.

Write a ground state electron configuration for these ions.

Remember that ions have a change in the total number of electrons (positive have lost electrons and negative have gained). Example: is  $1s^2 2s^2 2p^6$  . It has three extra electrons  $1s^2 2s^2 2p^6 3s^2 3p^2$  .  
11.  $O^{2-}$  12. Fe 13.  $CH_4$  14.  $Ni^{2+}$  15. 16.  $C_3H_8$  .

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Electron Configuration - Practice Problems KEY 1. Relate each of the following to Hog Hilton: Pauli Exclusion Principle - Each orbital can hold TWO electrons with opposite spins. Pigs Don't Like each other. Aufbau Principle - Electrons fill the lowest energy orbital's first. Pigs are Lazy. Hu...

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Writing electron configuration worksheet answer key. ... Electron Configurations Worksheet Practice Worksheets Electron Configuration Practices Worksheets Chemistry Worksheets . The electron configurations in this worksheet assume that lanthanum la is the first element in the 4f block and that actinium ac is the first element in the 5f block.

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Electron Configurations - Solutions Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different.

## **Electron Configuration Practice Worksheet 2003**

Write the abbreviated ground state electron configurations for the following: 8) helium  $1s^2$  (this one cannot be abbreviated) 9) nitrogen [He]  $2s^2 2p^3$  10) chlorine [Ne]  $3s^2 3p^5$  11) iron [Ar]  $4s^2 3d^6$  12) zinc [Ar]  $4s^2 3d^{10}$  13) barium [Xe]  $6s^2$  14) polonium

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[Xe] 6s24f145d106p4

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Electron Configurations Practice Questions 1. Predict the electron configurations of S, K, Ti, Sn. S:  $1s^2 2s^2 2p^6 3s^2 3p^4$  K:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$  Ti:  $1s^2 2s^2 2p^6 3s^2 3p^6 3d^2 4s^2$  ...

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Configuration Writing Practice Write a ground state electron configuration for each neutral atom. Ground state means that all of the lowest possible energy levels (up to the proper number of electrons for the element) are filled. 9. Cl 10. Hg 3 3&3p5 \ D. 1. Na 3. Sr 11.4 1-4 Write a ground state electron configuration for these ions.

### **KM 754e-20150323151140**

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