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In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made. Model 1 - Gases in a Nonflexible Container Experiment A (Adding more gas)

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Gavin Gutowsky Chemistry Blog: Gas Variables Pogil

Just remember the general gas law $PV=nRT$ or rearranging terms we see that $\frac{P_1V_1}{nT_1}=\frac{P_2V_2}{nT_2}$ where R is a constant. Therefore, $\frac{P_1V_1}{nT_1}=\frac{P_2V_2}{nT_2}$ this very general set of ratios relates all of the variables in state (1) to the same variables in state (2).

gas laws - The relationships between n , T , P , V for a ...

Oxygen gas . 1. The diagram in Model 1 illustrates a clipping of an aquatic plant in water. ... Provide specific evidence from Model 2 to support your answer. negatively - there are less bubbles with less light . Model 3 — Aquatic Plant Data ... variables that you think may change the outcome of the experiment, but since they are not being ...

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Answer Key. 1 - Patty Power. Which people are in the control group? Group B. What is the independent variable? New sauce. What is the dependent variable? Amount of gas. What should Mr. Krabs' conclusion be? The new sauce appears to work as it reduced the amount of gas. produced in 60% of the people tested.

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The volume of a gas depends on the pressure as well as the temperature of the gas. Therefore, a relation between these quantities and the mass of a gas gives valuable information about the physical nature of the system. Such a relationship is referred to as the equation of state.

Lab 10 - The Ideal Gas Law

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The gas volume is also directly proportional to temperature and number of moles. Therefore, V and T, V and n, and T and P are directly proportional to each other. Answers A, B, and D are then correct.