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Unit 12: Electrochemistry-Key Regents
Chemistry '14-'15 Mr. Murdoch Page 13
of 61 Key e. Solving for X, we get +6, so
the oxidation number for sulfur in the
polyatomic ion (SO_4^{2-}) is +6, and so
sulfur will be +6 in ZnSO_4 . *Check the
Periodic Table to see if sulfur has a +6
charge listed; it does!

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A) 12 moles B) 2 moles C) 3 moles D)6

moles 41.Given the balanced equation

representing a reaction: $2\text{Fe} + 3\text{Cu}^{2+} \rightarrow$

$2\text{Fe}^{3+} + 3\text{Cu}$ When the iron atoms lose

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six moles of electrons, how many moles of electrons are gained by the copper ions? A) 1 mole B) 6 moles C) 3 moles D) 9 moles 42. Given the balanced equation:

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Electrochemistry(redox) 68. Base your answer to the following question on the information and diagram below. The apparatus shown in the diagram consists of two inert platinum electrodes

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immersed in water. A small amount of an electrolyte, H_2SO_4 , must be added to the water for the ...

Redox Electrochemistry Regents Unit Review Answers

Oxidation: loss of electrons Reduction:
gain of electrons Oxidizing Agent:
substance reduced Reducing Agent:

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substance oxidized Anode: electrode where oxidation occurs (AN OX)
Cathode: electrode where reduction occurs (RED CAT) Salt Bridge: allows the flow of ions between the anode and cathode solutions to keep a balance of charge Cations flow toward the Cathode

Chemistry Regents Review: Redox &

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Redox And Electrochemistry Regents Short Answer Questions

Electrochemistry is the chemistry of electrochemical reactions; which deal with the relationship between electrical energy and chemical reactions.

Electrochemical reactions involve transfer of electrons and are essentially REDOX reactions. Displacement and

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REDOX Reactions Experiment 1:-
Displacement Reactions among Metals.
Procedure

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moles 41. Given the balanced equation representing a reaction: $2\text{Fe} + 3\text{Cu}^{2+} \rightarrow 2\text{Fe}^{3+} + 3\text{Cu}$ When the iron atoms lose six moles of electrons, how many moles of electrons are gained by the copper ions? A) 1 mole B) 6 moles C) 3 moles D) 9 moles

42. Given the balanced equation:

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PRACTICE PACKET: ELECTROCHEMISTRY
3 • All Free elements are zero • Group 1 is +1 • Group 2 is +2 • Fluorine is -1 • H is usually +1, and if so, will be first written in the chemical formula. • O is almost always -2, exceptions: H₂O₂ and OF₂ • In compounds using polyatomic ions, the first and middle

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elements all have + oxidation states. •
The sum of oxidation states in a ...

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