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How many sodium atoms are needed to react with  $1.25 \times 10^{24}$  molecules of water?

Ans:  $1.25 \times 10^{24}$   
atoms Na 1 mol H<sub>2</sub>O  
2 mol Na  $6.02 \times 10^{23}$   
atoms Na  $1.25 \times 10$   
molecules H<sub>2</sub>O x x x  
=  $6.02 \times 10^{23}$

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molecules  $\text{H}_2\text{O}$  2 mol

$\text{H}_2\text{O}$  1 mol Na 24

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Worksheet #1

continued 5. Hematite,  $\text{Fe}_2\text{O}_3$ , is an important ore of iron.

## **Stoichiometry Worksheet 1 Answer Key - Studylib**

Answer:  $8.75 \text{ g O}_2$  (1 mol  $\text{O}_2$  32.00 g  $\text{O}_2$ ) (2 mol  $\text{H}_2$  1 mol  $\text{O}_2$ ) (2.02 g  $\text{H}_2$  1 mol  $\text{H}_2$ ) = 1.10 g  $\text{H}_2$  (In your calculator:  $8.75 \div$

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$32.00 \times 2 \times 2.02 =$ )

## 13.3 Mass-Volume

Stoichiometry OR Molar

Mass gas @ STP Recall:

Avogadro's Molar

Volume is 22.4 L/mol

for a gas only at STP

Steps: 1) If given

grams, use MM as your

conversion factor to

get to moles of the

given

## **Chapter 13**

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Weight (FW)! • Sum of the atomic weights for the atoms in a chemical formula • So, the formula weight of calcium chloride,  $\text{CaCl}_2$ , would be Ca:  $1(40.1 \text{ amu}) + \text{Cl: } 2(35.5 \text{ amu})$  111.1 amu • These are generally reported for ionic compounds

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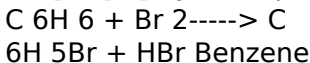
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## **Stoichiometry (Worksheet) - Chemistry**

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Stoichiometry Example



( $\text{C}_6\text{H}_6$ ) reacts with Bromine to produce bromobenzene ( $\text{C}_6\text{H}_5\text{Br}$ ) and hydrobromic acid. If 30. g of benzene reacts with 65 g of bromine and produces 56.7 g of bromobenzene, what is the percent yield of the reaction? 30.g 65 g 56.7 g

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## Chapter 3

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basic introduction into stoichiometry.

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Create a new spreadsheet in your workbook and name it 'Practice 1'. In this spreadsheet, create a table using formulas in Excel that can calculate the grams of Al (s) and Cl<sub>2</sub> (g) required to produce a desired amount of

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$\text{AlCl}_3$  (s) according to  
the reaction  $2 \text{Al}(s) + 3 \text{Cl}_2 (g) \rightarrow 2 \text{AlCl}_3 (s)$   
Set up your table in a  
similar manner to that  
shown in Worked  
Example #1.

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