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Texas, Dallas, TX, and  
Jane D. Smith,

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improve students'  
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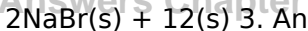
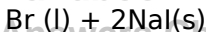
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Note: Not all of the reactions will occur. For those that do not, write no reaction. 1. A piece of copper is dropped into a container of water. No reaction 2.

Liquid bromine is

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added to a container of sodium iodide crystals.



3. An aluminum strip is immersed in a solution of silver nitrate.

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Free elements, no matter how complex the molecule, have an

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oxidation number

(valence or charge)

equal to zero. O, Fe.

T ~ He < /1 1 ~ ~ L-td.

8. The following are  
diatomic or polyatomic  
elements in nature  
which must be  
committed to memory.

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come up with the  
money for you not by  
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Reactions(See  
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28, Practice-Pg. 49 •

Formation of a Gas -

Pg. 50, Practice-Pg. 51

• Acid Base

Neutralization

Reactions- Pg. 53,

Practice-Pg.54

Weak acids and bases do not

break apart and are in

net ionic equation. •

Synthesis and

Decomposition

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Single Replacement

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## Quiz on all Chemical Reactions - Max Study

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SnCO tin(II) carbonate  
sodium hydrogen  
carbonate NaHCO 1. 2.  
3. 4. 5. 6. 8. 9. 10. 11.  
vanadium(V) oxide H<sub>2</sub>O  
dihydrogen monoxide  
(NH)<sub>4</sub>CO ammonium  
oxalate 4 224  
polonium(VI)  
thiocyanate Po(SCN)<sub>6</sub>  
tetraphosphorus



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decaoxide P O 4 10 12.  
13. 15. 16. 17. 18. 19.  
manganese(VII) oxide  
copper(II) dihydrogen  
phosphate

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George R. Hauge, Jane  
D. Smith, June 2001,  
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Scientific, Inc.

Correlating questions

(ii) from: Rodney

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Schmitz (2007) Moore  
County Schools, NC

Reaction Prediction - 5

For each of the

following eight

reactions, in part (i)

write a BALANCED

equation and in part (ii)

answer the question

about the reaction.

## **KEY to 5 Reaction Prediction -**

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Determine the

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oxidation number of  
the underlined  
element:  $\text{KMnO}_4$  Since  
K is an alkali metal, its  
charge must be  $1+$   
Oxygen is  $2-$  but there  
are four of them,  
therefore, 4 times  $2-$   
equals  $8-$  If  $1+$  and  
 $8-$  are added Ultimate  
Chemical Equation  
Answers

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Key**

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Determine the  
oxidation number of  
the underlined  
element:  $\text{KMnO}_4$  Since  
K is an alkali metal, its  
charge must be  $1+$   
Oxygen is  $2-$  but there  
are four of them,  
therefore, 4 times  $2-$   
equals  $8-$  If  $1+$  and  
 $8-$  are added The  
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